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Atomic Structure. The

atom consists of three

components: Protons,

Neutrons, and

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Number. The order of

an element in

Mendeleev's table of the

elements; Often

represented by the

symbol Z and it dictates

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in the nucleus. Atomic
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summarize our current
understanding of atomic

structure. include the

two parts, where the

three particles are

located, and where the

mass and volume in an

atom mainly exist. the

atoms is made up of a

nucleus and electron

cloud. in the nucleus it is

made up of subatomic

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II) Structure of the
atom. Subatomic
Particles (sizes, charges,
locations) Protons.
Electrons. Neutrons.
Atomic Number vs.
Mass Number: Know
how to read an element
square off of Table.

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Atomic Number: know what it is and how to find it. Mass Number: know what it is and how to find it. How do you know how many electrons there are in an atom?

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Chapter 4: Atomic

Structure 4.1 Studying

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Atomic Theory John

Dalton studied the behavior of gases in air, and concluded that gas consists of individual

particles. Evidence for Atoms: No matter how large or small the sample, the ratio of the masses of the element in the compounds is always the same. (Compounds have fixed composition.)

Dalton proposed the

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theory that all matter is made up of individual particles called atoms, which cannot be divided. Main ...

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of the wavelength or frequency of the radiation. Historically, spectroscopy originated as the study of the wavelength dependence of the absorption by gas phase matter of visible light dispersed by a prism. Matter waves and acoustic waves can also be considered forms of radiative energy, and

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atoms, calculate the
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valence electrons in an
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